ABOVE & BEYOND

Purpose …

- Y11 Intervention will be known as '**Above & Beyond'.**
- The idea is that we will support you in going 'Above and Beyond' and if you do, you will achieve your potential!
- We expect you to attend after school Above & Beyond sessions (Period 7's) - **Compulsory.** This is to support you with your learning and we will ensure this is a positive learning experience. **We are in this together!**
- Morning intervention - Form tutor will let you know where to go for this
- Above and Beyond Success Card in your planners - stamped

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ABOVE & BEYOND

Above & Beyond Success Card Prizes…

Completed Cards	Eligible Reward
1 card	Chocolate bar
3 cards	Academic Stationary
5 cards	Weekly Skip the Queue Pass
6 cards	Free Breakfast
7 cards	Free Lunch
9 cards	Water Bottle
10 cards	x2 Cinema Tickets
12 cards	GOLDEN PROM TICKET

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ABOVE & BEYOND

“Learn, Revise, Test Yourself”

Thursday 3rd October 2024

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ABOVE & BEYOND

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Additional Notes

Please can we have this page in your planners out for the duration of this sessions and take notes on useful information

Above & Beyond Timetable (HT1)…

Date	Focus	Staff
Tues 1st Oct	**"Getting Revision Ready"**	*Mr Kenny*
Thurs 3rd Oct	**'Learn, Revise, Test Yourself'**	*Mr Fuller*
Tues 8th Oct	**Condensing your notes**	*Mrs Holme*
Wed 9th Oct	**Careers Talk - Engineering**	*MFT*
Thurs 10th Oct	**Memory Techniques**	*Mr Loftus*
Tues 15th Oct	**Past Papers & Exam Q's**	*Mrs Dawson*
Wed 16th Oct	**Careers Talk - NHS Senior Manager**	*NHS*
Thurs 17th Oct	**Flash Cards & Mind Maps**	*Mr Floweth*

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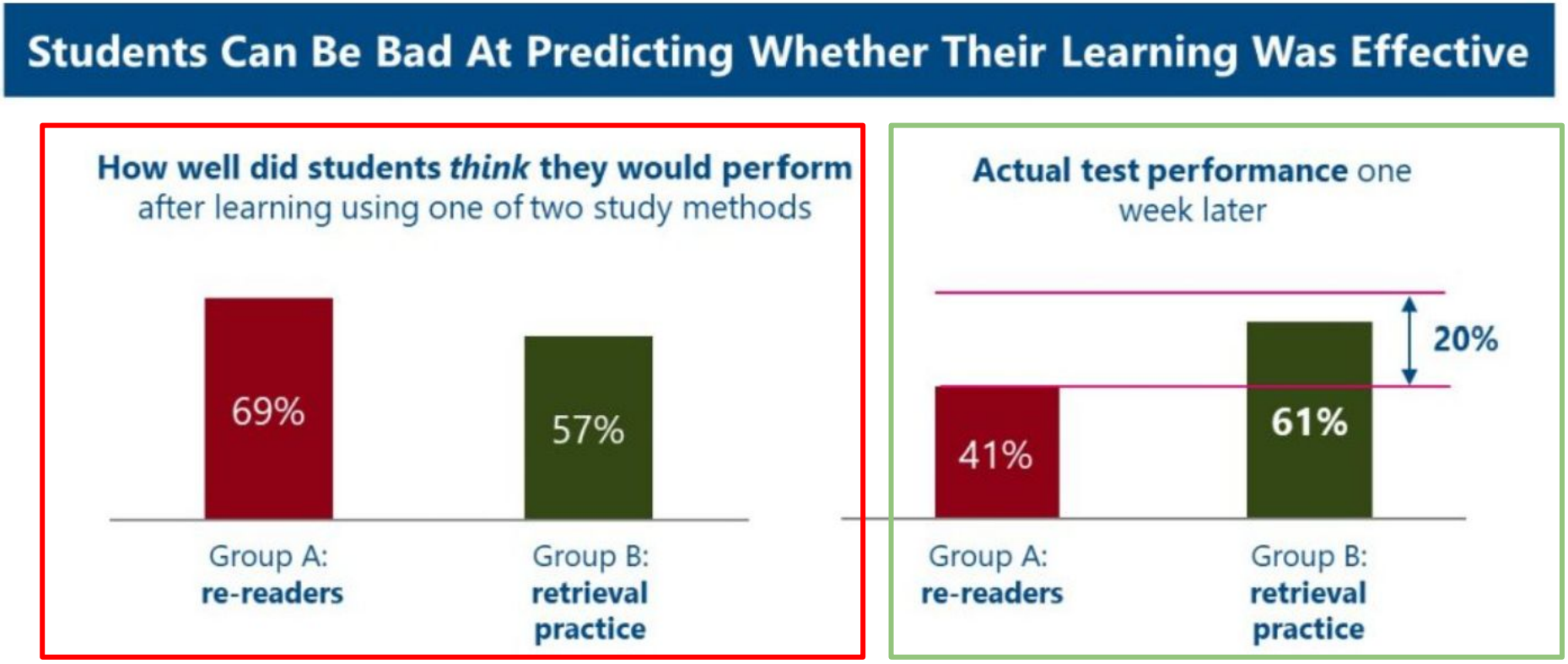
Learn, Revise & Test Yourself



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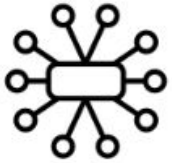
Learn, Revise, Test Yourself

Strategy 1

Flash Cards

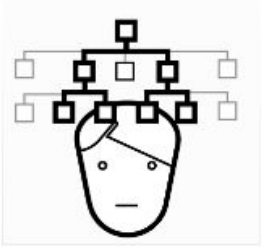
Strategy 2

Mind Maps



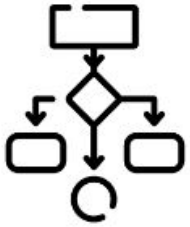
Strategy 3

Mnemonics



Strategy 4

Flow Charts



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ABOVE & BEYOND

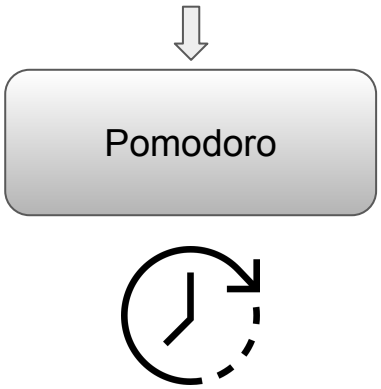
Pomodoro Technique

WHAT IS THE POMODORO TECHNIQUE?

A method for staying focused and mentally fresh

STEP 1		Pick a task
STEP 2		Set a 25-minute timer
STEP 3		Work on your task until the time is up
STEP 4		Take a 5 minute break
STEP 5		Every 4 pomodoros, take a longer 15-30 minute break

Strategy



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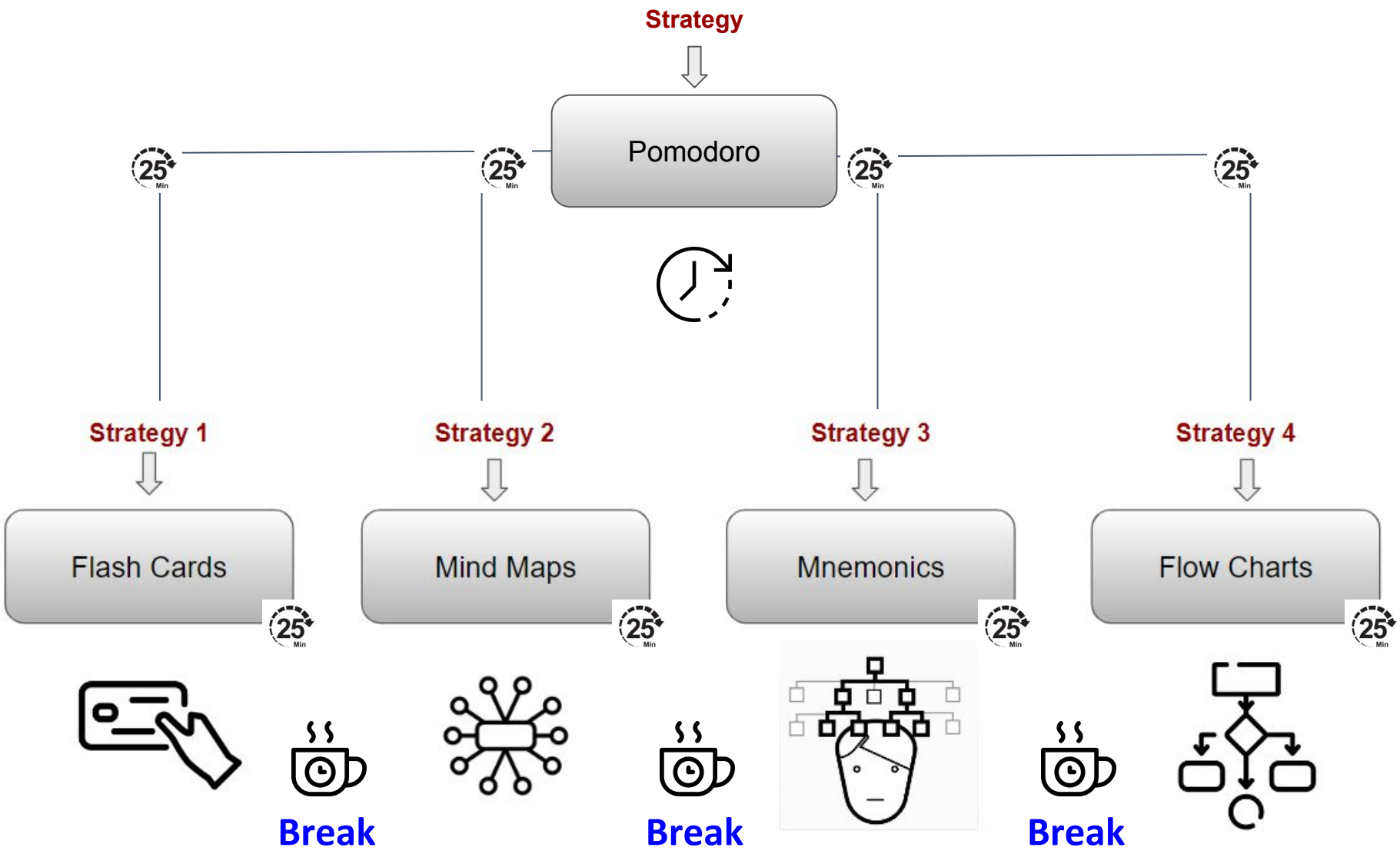
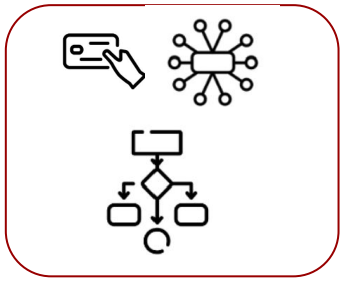
Pomodoro Technique



Pomodoro in Action

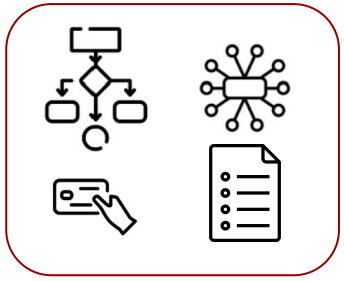
Work 25 Min — Maths



Test Yourself
Exam Questions & Check the Mark Scheme

Break

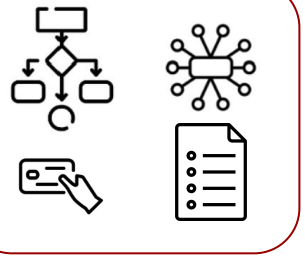
Work 25 Min — Science



Test Yourself
Exam Questions & Check the Mark Scheme

Break

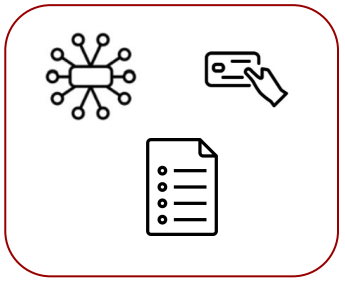
Work 25 Min — History



Test Yourself
Exam Questions & Check the Mark Scheme

Break

Work 25 Min — English



Test Yourself
Exam Questions & Check the Mark Scheme

Revision Steps to Success

Step 1: Decide on the **subject** your going to revise

Step 2: Select the **revision strategy** you are going to use to support your revision

Step 3: Creating a mind map, flash cards, flow chart on the **knowledge** needed using the revision guide

Step 4: Test yourself by re-creating the mind map or timeline from memory without the revision guide

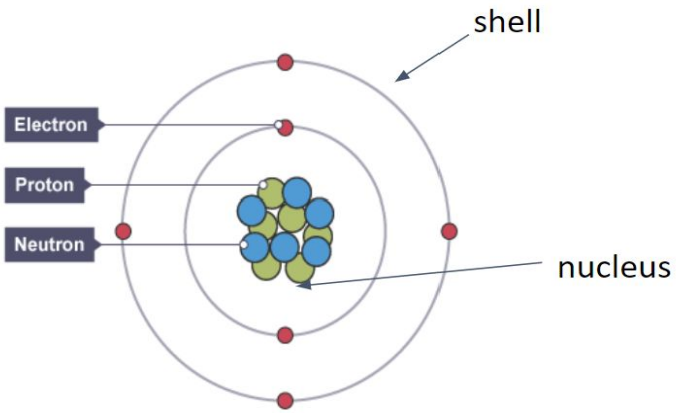
Step 5: Check & add in key information you missed in a different colour, then test yourself again

Task 1: Atomic Structure (4 mins)

Atomic Structure

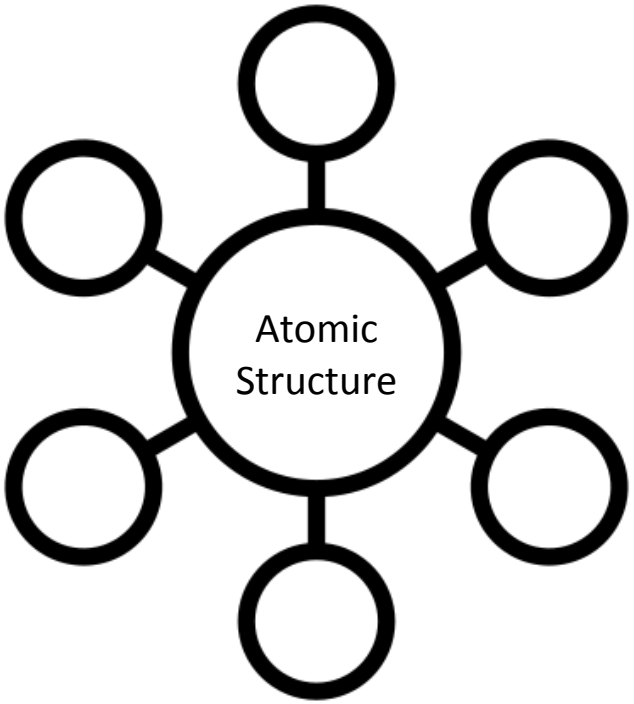
An **atom** consists of a central nucleus, which:

- Contains **protons** and **neutrons**
- Is surrounded by electrons in shells



Particle	Relative Charge	Relative Mass
Proton	+1	1
Neutron	0	1
Electron	-1	$\frac{1}{1836}$

Mind Map

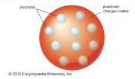


Step 1: Check what you have not included and add to it

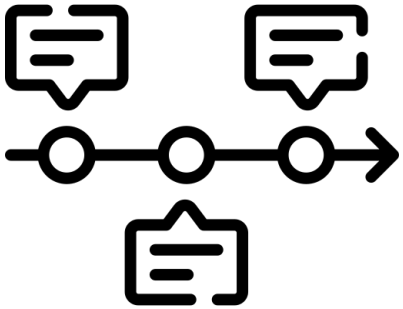
Green Pen

Step 2: Blank paper and repeat to test yourself again

Task 2: Atomic Structure

When	Who	What
1803	Dalton	Solid atom model: all atoms of an element are identical; different elements have different atoms
1897	Thomson	Discovers the electron
1904	Thomson	Plum pudding model: atoms are spheres of positive charge with negative electrons dotted around inside 
1911	Rutherford	Solar system model: atoms have a positive nucleus surrounded by negative electrons in orbits
1913	Bohr	Electron shell model: electrons occupy shells or energy levels around the nucleus
1918	Rutherford	Discovers the proton
1932	Chadwick	Discovers the neutron

Timeline or Flowchart



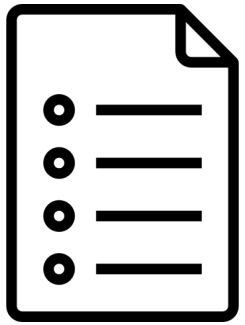
Step 1: Check what you have not included and add to it

Green Pen

Step 2: Blank paper and repeat to test yourself again

Task 3: Atomic Structure (Exam Questions)

Exam Questions



Q1.

The scientist John Dalton lived over 200 years ago.

John Dalton suggested an early model of atoms.

When Dalton first described atoms he said that

- all elements are made of atoms
- atoms are not formed of any smaller particles
- all atoms of the same element are identical.

Give two differences between Dalton's model of atoms and today's model of atoms.

(2)

1 .....

.....

.....

2 .....

.....

.....

(Total for question = 2 marks)

Q2.

Task 3: Atomic Structure (Exam Questions)

(a) Atoms contain protons, neutrons and electrons.

Complete the table to show the relative mass and relative charge of each particle and its position in an atom.

(3)

	relative mass	relative charge	position in atom
proton		+1	
neutron	1		in nucleus
electron			

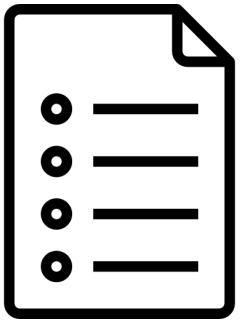
(b) Complete the sentence by putting a cross (☒) in the box next to your answer.

An atom of an element **always** contains

(1)

- ☒ **A** more protons than neutrons
- ☒ **B** equal numbers of protons and neutrons
- ☒ **C** more electrons than protons
- ☒ **D** equal numbers of protons and electrons

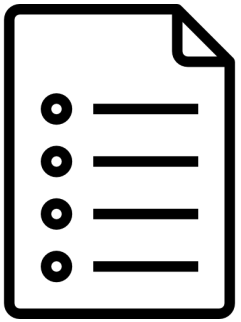
Exam Questions



Task 3: Atomic Structure (Exam Mark Scheme)

Green Pen

Exam Questions



Mark Scheme

Q1.

Question number	Answer	Additional guidance	Mark
	Any two from (in modern model) • atoms are formed of sub-atomic particles (1) • atoms have a nucleus (1) • atoms contain protons (1) • atoms contain neutrons (1) • atoms contain (shells of) electrons (1) • atoms of same element can have different numbers of neutrons / isotopes exist (1)	allow (for Dalton's model) atoms are indivisible ignore statements that are simply the negative of those in the question reject each comparison with 'plum pudding model'	(2) AO1

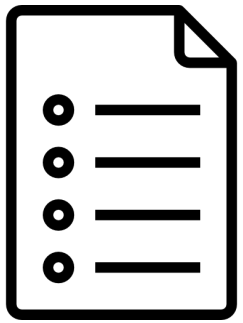
Task 3: Atomic Structure (Exam Mark Scheme)

Green Pen

Question Number	Answers	Acceptable Answers	Mark
(a)	<table><tr><th></th><th>relative mass</th><th>relative charge</th><th>position in atom</th></tr><tr><td>proton</td><td>1</td><td>(+1)</td><td>in nucleus</td></tr><tr><td>neutron</td><td>(1)</td><td>0</td><td>(in nucleus)</td></tr><tr><td>electron</td><td>1/1837</td><td>-1</td><td>in shells</td></tr></table> all 6 correct (3) 4 or 5 correct (2) 2 or 3 correct (1)	ignore units reject relative mass of proton: +1/1+ for relative mass of electron: anything smaller than 1/1500/0.00067 (almost) 0/negligible/very small for relative charge on neutron: none/no charge/neutral for position of electron in an atom: in orbits / orbitals / energy levels / around the nucleus /outside the nucleus ignore rings ignore inner/outer	**(3)**

Question Number	Answers	Acceptable Answers	Mark
(b)	D equal numbers of protons and electrons		**(1)**

Exam Questions

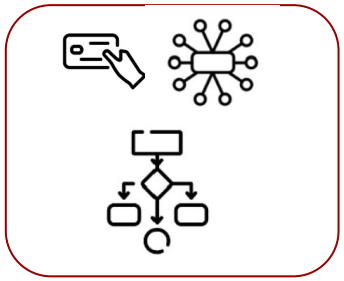


BREAK (5 mins)

Pomodoro in Action

Pt 1 of the Pomodoro complete

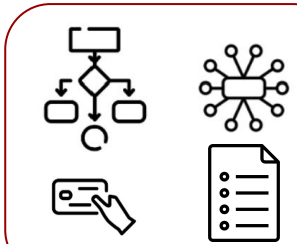
Work 25 Min — Maths



Test Yourself
Exam Questions & Check the Mark Scheme

Break

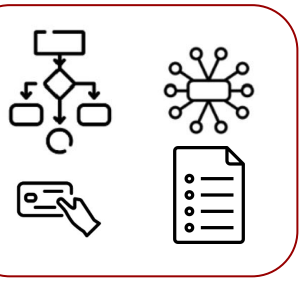
Work 25 Min — Science



Test Yourself
Exam Questions & Check the Mark Scheme

Break

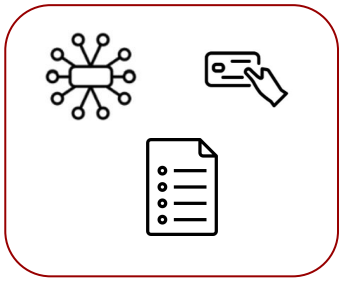
Work 25 Min — History



Test Yourself
Exam Questions & Check the Mark Scheme

Break

Work 25 Min — English



Test Yourself
Exam Questions & Check the Mark Scheme

ABOVE & BEYOND